

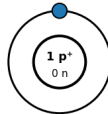
# Atome der 1. bis 3. Periode - Schalenmodell

Manggers  
Lernmaterialien

## 1. Periode

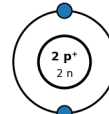
1 Schale

Wasserstoff (H)



Elektronen: 1

Helium (He)

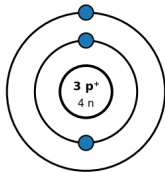


Elektronen: 2

## 2. Periode

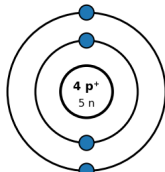
2 Schalen

Lithium (Li)



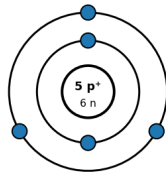
Elektronen: 2 | 1

Beryllium (Be)



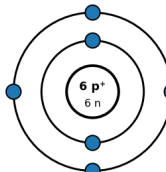
Elektronen: 2 | 2

Bor (B)



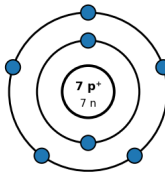
Elektronen: 2 | 3

Kohlenstoff (C)



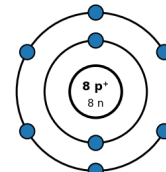
Elektronen: 2 | 4

Stickstoff (N)



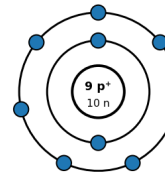
Elektronen: 2 | 5

Sauerstoff (O)



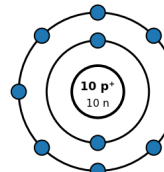
Elektronen: 2 | 6

Fluor (F)



Elektronen: 2 | 7

Neon (Ne)

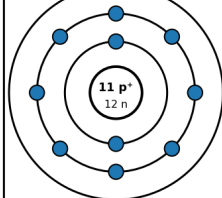


Elektronen: 2 | 8

## 3. Periode

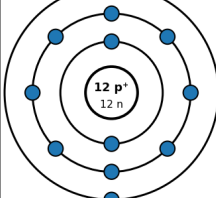
3 Schalen

Natrium (Na)



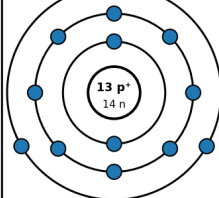
Elektronen: 2 | 8 | 1

Magnesium (Mg)



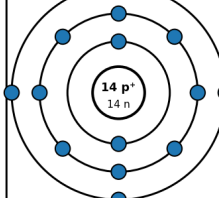
Elektronen: 2 | 8 | 2

Aluminium (Al)



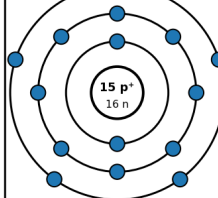
Elektronen: 2 | 8 | 3

Silicium (Si)



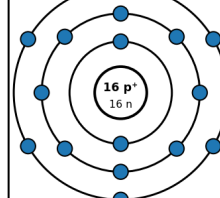
Elektronen: 2 | 8 | 4

Phosphor (P)



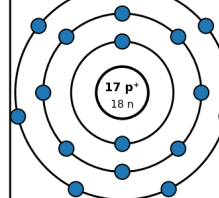
Elektronen: 2 | 8 | 5

Schwefel (S)



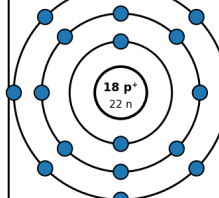
Elektronen: 2 | 8 | 6

Chlor (Cl)



Elektronen: 2 | 8 | 7

Argon (Ar)



Elektronen: 2 | 8 | 8

**Füllregel: zuerst 1. Schale (max. 2 e<sup>-</sup>), danach 2. Schale (max. 8 e<sup>-</sup>), danach 3. Schale.**

Elektronenverteilungen: H 1 · He 2 · Li 2|1 · Be 2|2 · B 2|3 · C 2|4 · N 2|5 · O 2|6 · F 2|7 · Ne 2|8 · Na 2|8|1 · Mg 2|8|2 · Al 2|8|3 · Si 2|8|4 · P 2|8|5 · S 2|8|6 · Cl 2|8|7 · Ar 2|8|8